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$[\mu_2-O, O', O'', OO''']$ -Bis
(1,2-dithiooxalato-*S,S'*)nickel(II)]
bis[-*O, O'*-bis(1,2-dithiooxalato-*S,S'*)-
nickel(II)pentaquaholmium(III)]hydrate,
 $[Ho_2Ni_3(dto)_6(H_2O)_{10}]$

Suggested citation referring to the original publication:

Molbank 2016(2)

DOI <http://dx.doi.org/10.3390/M895>

ISSN 1422-8599

Postprint archived at the Institutional Repository of the Potsdam University in:

Postprints der Universität Potsdam

Mathematisch-Naturwissenschaftliche Reihe ; 360

ISSN 1866-8372

<http://nbn-resolving.de/urn:nbn:de:kobv:517-opus4-400598>

Short Note

$[\mu_2\text{-O}, \text{O}', \text{O}'', \text{OO}''']\text{-Bis(1,2-dithiooxalato-}S, S')\text{nickel(II)}]$ $\text{bis}[\text{-O}, \text{O}'\text{-bis(1,2-dithiooxalato-}S, S')\text{-nickel(II)}]$ $\text{pentaquaholmium(III)]hydrate, [Ho}_2\text{Ni}_3(\text{dto})_6(\text{H}_2\text{O})_{10}]$

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Academic Editor: Norbert Haider

Received: 9 February 2016; Accepted: 5 April 2016; Published: 12 April 2016

Abstract: Planar bis(1,2-dithiooxalato)nickelate(II), $[\text{Ni}(\text{dto})]^{2-}$ reacts in aqueous solutions with lanthanide ions (Ln^{3+}) to form pentanuclear, hetero-bimetallic complexes of the general composition $[\{\text{Ln}(\text{H}_2\text{O})_n\}_2\{\text{Ni}(\text{dto})_2\}_3] \cdot x\text{H}_2\text{O}$. ($n = 4$ or 5 ; $x = 9\text{--}12$). The complex $[\{\text{Ho}(\text{H}_2\text{O})_5\}_2\{\text{Ni}(\text{dto})_2\}_3] \cdot 10\text{H}_2\text{O}$, Ho_2Ni_3 , was synthesized and characterized by single crystal X-ray structure analysis and powder diffraction. The Ho_2Ni_3 complex crystallizes as monoclinic crystals in the space group $P2_1/c$. The channels and cavities, appearing in the crystal packing of the complex molecules, are occupied by a varying amount of non-coordinated water molecules.

Keywords: lanthanides; holmium(III); 1,2-dithiooxalate; crystal structure; nickel(II)

1. Introduction

During the last decades, molecular hetero-bimetallic 3d–4f complexes with transition metals and trivalent lanthanide ions have gathered increasing interest due to their magnetic properties [1–17]. Such complexes are promising precursors for, e.g., catalysts, magnetic materials, luminescent materials and new molecular devices. Since lanthanide ions provide a large angular momentum and the f–f transition is less influenced by the ligand field compared to the spin-orbit coupling, magnetic behavior different from the transition metals has been observed [16]. When transition metals are bridged to the lanthanide ions by small bridging ligands 3d–4f complexes with interesting magnetic properties are formed, e.g., as single-molecule magnets [18]. The design and construction of extended hetero-nuclear 3d–4f complexes with a discrete multinuclear aggregation is still a challenge for coordination chemists. A series of pentanuclear, hetero-bimetallic 3d–4f complexes of the general composition $[\{\text{Ln}(\text{H}_2\text{O})_n\}_2\{\text{Ni}(\text{dto})_2\}_3] \cdot x\text{H}_2\text{O}$ ($n = 4$ or 5 ; $x = 9\text{--}12$) was first described by Trombè, Gleizes and Galy [19,20]. Dependent on the decreasing ionic radii of the lanthanide ions, two types of structures were observed: a monoclinic structure for the larger lanthanide ions (La–Dy) and a triclinic structure for the smaller lanthanide ions (Er, Yb). The analogue Ho_2Ni_3 complex is the borderline case between both structural forms and its structural type was unclear. Here we can report the structure of this complex $[\{\text{Ho}(\text{H}_2\text{O})_5\}_2\{\text{Ni}(\text{dto})_2\}_3] \cdot 10\text{H}_2\text{O}$, Ho_2Ni_3 . The here reported complex reported here crystallizes in the monoclinic space group $P2_1/c$. The structural information of the Ho_2Ni_3 complex is new in this series to the best of our knowledge.

2. Results and Discussion

2.1. X-ray Structures

A single crystal of the complex Ho_2Ni_3 was measured by X-ray diffraction. Table S1 summarizes the crystallographic data and refinement parameters for this complex. The Ho_2Ni_3 complex crystallizes

in the monoclinic space group $P2_1/c$. In the crystals, two different forms of enclosed water molecules can be observed: defined coordinated water at the lanthanide ion and non-stoichiometric amounts of inserted water in the cavities and between the layers of the complex molecules.

Figure 1 shows the molecular structure of the pentanuclear heterobimetallic Ho_2Ni_3 complex with the characteristic z-shape of this type of molecule [21–24]. The two holmium centers are bridged by a nearly planar bis(1,2-dithiooxalato)nickelate(II) unit. Two peripheric bis(1,2-dithiooxalato)nickelate(II) moieties are coordinated only bidentate in a non-bridging mode. The structure of the complex is close to planarity with a little twisting, due to the non-symmetric coordination geometry of the lanthanide ions (see Figure 2). The coordination spheres of the lanthanide ions are completed by five coordinated water ligands at each lanthanide ion, resulting in a coordination number of nine. The nine donor atoms in the coordination sphere of the lanthanide ion occupy the corners of a tricapped trigonal prism. The molecule is centrosymmetric with a crystallographic symmetry center at the Ni1 atom.

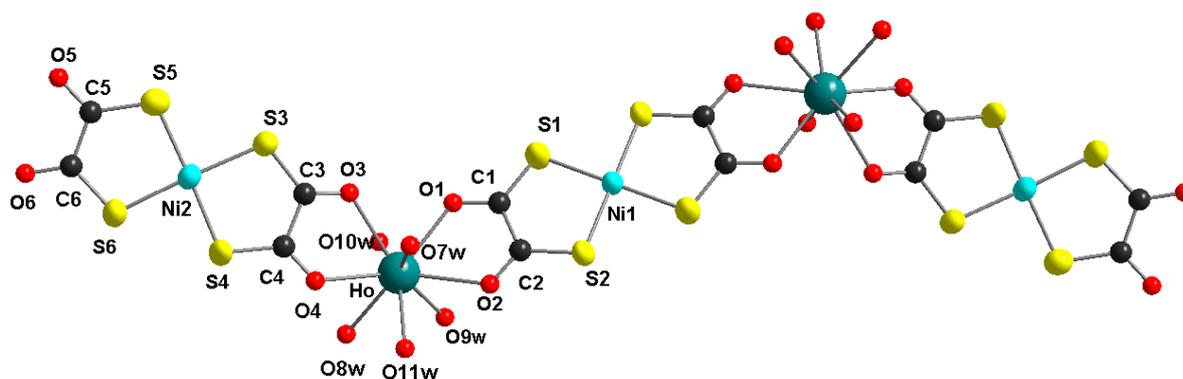


Figure 1. Molecular structure of the Ho_2Ni_3 complex, view along the crystallographic a -axis (non-coordinated water molecules and hydrogen atoms are omitted for clarity).

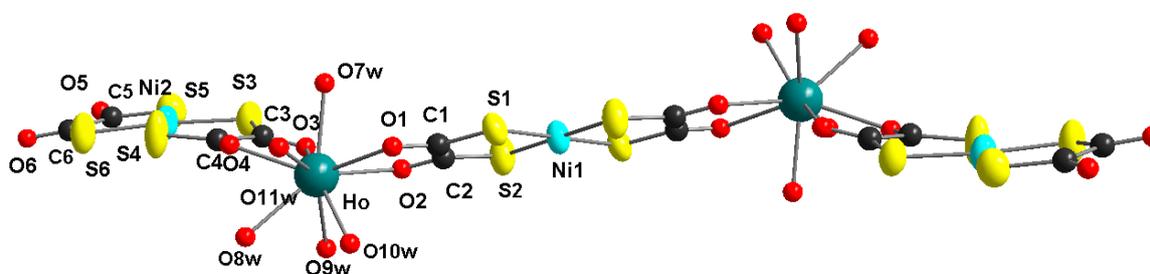


Figure 2. Molecular structure of the Ho_2Ni_3 complex, view along the crystallographic b -axis (non-coordinated water molecules and hydrogen atoms are omitted for clarity).

The complex is almost isostructural to the already reported analogue Eu_2Ni_3 , Nd_2Ni_3 , Ce_2Ni_3 , and Dy_2Ni_3 complexes [21–24]. The deviations in the structures are only due to small differences of the ion radii of the individual lanthanide ions. The crystal packing of the molecules contains channels and cavities, which are filled by non-coordinated water molecules. Whereas the coordinated water molecules are in fixed positions, the non-coordinated water molecules are strongly disordered. The slightly varying number of non-coordinated water molecules could be calculated (estimated) by thermoanalytical investigations and was determined with approximately nine to 12 inserted water molecules per formula unit [22]. This inserted water is very flexible and already partly released at laboratory conditions with dry air, making it difficult to determine their exact number. This also applies to the Ho_2Ni_3 complex in this series. Therefore, the crystal was completely embedded in perfluoropolyalkyl ether to protect it against decay during the X-ray crystal structure analysis. This effect is also responsible for the relatively high remaining electron density for this type of complex.

2.2. IR Spectroscopy

All these complexes show intense water signals around 3300 cm^{-1} in their IR absorption spectra (see Figure S2). Furthermore, very broad carbonyl vibrations (between 1600 cm^{-1} and 1400 cm^{-1}) can be observed. This broadening results from the overlapping of the different types of coordinated and non-coordinated carbonyl groups present in the complex: non-coordinated terminal carbonyl groups of the peripheric 1,2-dithiooxalato ligands, which are found at approximately 1600 cm^{-1} , comparable to the red form of the mononuclear $\text{K}_2[\text{Ni}(\text{dto})_2]$ complex [25], combine with the vibration of the carbonyl groups of the bridging dto ligands coordinated by the lanthanide ions. Also, the formation of intermolecular hydrogen bonds towards some of the carbonyl groups contributes to the broadening of the CO vibrations in the spectra. The resulting absorption bands of carbonyl vibrations are found at around 1495 cm^{-1} . As already mentioned, the complexes of this series crystallize in two different crystal systems depending on the decreasing lanthanide radii, and this can also be recognized by IR spectroscopy. The comparison of vibration bands of $\nu(\text{C-C-S})$, $\nu(\text{S-C})$ and $\nu(\text{Ni-S})$ exhibits a shift of the absorptions in the range of 10 cm^{-1} if the crystal system changes from monoclinic to triclinic. These shifts are attributed to stronger bonding of the lanthanide ions and the dto ligands.

3. Experimental Section

3.1. General Methods, Analytical and Physical Measurements

All infrared spectra were recorded on a Perkin Elmer 16PC FT-IR-spectrometer in a range between 400 cm^{-1} and 4000 cm^{-1} using KBr pellets. The elemental analyses (C, H, S) were determined by a Vario EL III CHNS from elementar Analysensysteme GmbH (Hanau, Germany). The magnetic susceptibility was measured by a magnetic susceptibility balance MSB-Auto by Sherwood Scientific Ltd. at room temperature. XRD measurements were conducted with a Bruker AXS (Siemens) D5005 diffractometer using $\text{CuK}\alpha$ radiation ($\lambda = 1.540598\text{ \AA}$) (see Figure S1).

X-ray structure of the presented complex was collected using a STOE Image Plate Diffraction System IPDS-2 at 210 K with graphite-monochromatized $\text{MoK}\alpha$ radiation. The reflection data were corrected by an absorption correction using the program X-Area [26]. The structures were solved with SHELXS-97 [27] using direct methods and refined with SHELXL-2014 [28]. The nonhydrogen atoms were refined anisotropically, with the exception of one non-coordinated oxygen atom of a crystal water molecule with an occupation factor of 0.25 in the structure. The hydrogen atoms of the water molecules could not be found. CCDC 1450593 contains the supplementary crystallographic data for this article and can be obtained free of charge from The Cambridge Crystallographic Data Centre via http://www.ccdc.cam.ac.uk/data_request/cif (or from the CCDC, 12 Union Road, Cambridge CB2 1EZ, UK; Fax: +44 1223 336033; E-mail: deposit@ccdc.cam.ac.uk).

3.2. Synthesis of $[\text{Ho}_2\text{Ni}_3(\text{dto})_6(\text{H}_2\text{O})_{10}] \cdot 10\text{H}_2\text{O}$

The ligand potassium 1,2-dithiooxalate, K_2dto , was synthesized by sulfhydrolysis of diphenyl oxalate according to the procedure of Matz and Mattes [29] modified by Wenzel *et al.* [30]. The complex Ho_2Ni_3 was prepared according to the general procedure described by Trombe *et al.* [20].

A solution of K_2dto (1 mmol; 200.2 mg) in 5 mL of distilled H_2O was added to a stirred solution of $\text{NiCl}_2 \cdot 6\text{H}_2\text{O}$ (0.5 mmol; 119.7 mg) in 4 mL of distilled H_2O . The resulting dark-violet solution was heated to $50\text{ }^\circ\text{C}$. A warm solution ($50\text{ }^\circ\text{C}$) of $\text{HoCl}_3 \cdot x\text{H}_2\text{O}$ (0.33 mmol) in 4 mL water was added drop wise to the stirred $[\text{Ni}(\text{dto})_2]^{2-}$ solution. The resulting reaction mixture was continuously stirred at $50\text{ }^\circ\text{C}$ for additional 15 min, than slowly cooled down to room temperature. The dark-violet crystalline precipitate was filtered off, washed with a small amount of distilled water and dried at $80\text{ }^\circ\text{C}$.

K_2dto : $\text{C}_2\text{O}_2\text{S}_2\text{K}_2$ ($M = 198.35\text{ g/mol}$). Elementary analysis (EA) measured (calculated): C 11.95 (12.11); S 32.08 (32.33) %. IR (KBr): 1530, 1514 ($\nu_{\text{C-O}}$), 1113 ($\nu_{\text{C-C-S}}$), 879 ($\nu_{\text{C-S}}$) cm^{-1} .

Ho₂Ni₃, 1: C₁₂H₂₂O₂₃S₁₂Ni₃Ho₂ (M = 1425.01 g/mol). EA meas. (calc.): C 10.05 (10.11); H 1.55 (1.56); S 26.58 (27.00) %. IR (KBr), Figure S2: 3166 (ν_{OH}), 1490 ($\nu_{\text{C-O}}$), 1138 ($\nu_{\text{C-C-S}}$), 990, ($\nu_{\text{C-S}}$), 458 ($\nu_{\text{Ni-S}}$, δ_{ring}) cm⁻¹. μ_{eff} meas. (calc.): 14.63 (14.99) B.M.

Supplementary Materials: The following are available online at <http://www.mdpi.com/1422-8599/2016/2/M895>, Table S1: Crystallographic data and refinement parameters for the complex Ho₂Ni₃, Figure S1: Experimental X-ray powder patterns for Ho₂Ni₃ (1), Figure S2: IR absorption spectrum (KBr) of Ho₂Ni₃.

Acknowledgments: Christina Günter, University of Potsdam (Germany), Faculty of earth and environmental sciences for the powder diffraction measurements. Matthias Siebold, formerly University of Potsdam (Germany), Insitute of Chemistry for additional synthetic experiments.

Author Contributions: Jana König, Peter Strauch: writing of manuscript, analytical interpretation, literature search; Jana König: synthetic work; Alexandra Kelling, Uwe Schilde: X-ray crystallography.

Conflicts of Interest: The authors declare no conflict of interest.

Abbreviations

The following abbreviations are used in this manuscript:

dto 1,2-dithiooxalate

Ln lanthanide

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